

# Synthesis and photophysical properties of new Eu(III) complex

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## Abstract

In this study, the new Eu(III) complex was hydrothermally synthesized using oxygen donor ligand and characterized by single crystal X-ray diffraction, elemental analysis, solid UV-Vis spectroscopy, FT-IR and photoluminescence properties. The crystal structure analysis of complex  $[C_7H_9EuO_9S](I)$  indicates that the central lanthanide ions are coordinated by eight oxygen atoms in which six oxygen atoms from four  $ssa^-$  anionic ligands and two oxygen atoms from a coordinated water molecules. Complex **I** bridged by oxygen atoms from carboxyl and sulfonate groups of the  $ssa^-$  ligands are formed in 1D chains. The 1D chains are linked together with ligand  $ssa^-$  to form a 2D wave-like layers. All these wave-like layers are further interlinked via strong hydrogen bonding interaction resulting in 3D architecture. Additionally, solid state photoluminescence properties of complex **I** have been investigated in room temperature. The indirect energy transfer mechanism via an Antenna effect in complex **I** has been investigated in detail.

**Keywords:** Eu(III) complex, crystal structure, antenna effect.

## Yeni Eu(III) kompleksinin sentezlenmesi ve fotofiziksel özellikleri

### Özet

Bu çalışmada, oksijen donör ligandı kullanılarak yeni Eu(III) içeren lantanit kompleksi hidrotermal yöntem ile sentezlenmiş ve tek kristal X-ışını kırınımı, element analizi, katı UV-Vis spektroskopisi, FT-IR ve fotolüminesans özellikleri ile karakterize edilmiştir. Eu(III) kompleksinin  $[C_7H_9EuO_9S](I)$  kristal yapısı sekiz oksijen atomu ile koordine olduğunu göstermektedir. Burada altı oksijen atomu  $ssa^-$  anyonik ligandından ve iki oksijen atomu ise koordine su moleküllerinden gelmektedir. Kompleks **I**,  $ssa^-$  ligandının karboksil ve sülfonat grupları ile köprü kurarak bir-boyutlu zincir oluşturmaktadır. Bir boyutlu zincirler  $ssa^-$  ligandları ile birlikte iki-boyutlu dalga benzeri tabakalar oluşturmaktadır. Tüm bu dalga benzeri tabakalar güçlü hidrojen bağı etkileşimleri üç-boyutlu mimari ile sonuçlanmıştır. Buna ek olarak, kompleks (**I**)' in katı durum fotolüminesans özellikleri oda sıcaklığında araştırılmıştır. Kompleks için anten etkisi ile dolaylı enerji transfer mekanizması detaylı olarak araştırılmıştır.

**Anahtar kelimeler:** Eu(III) kompleksi, kristal yapı, anten etkisi.

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## 1. Introduction

In recent years, the construction of coordination polymers has been a field of rapid growth not only for their intriguing topologies and diverse structures but also for their interesting chemical and physical properties, such as magnetism, luminescence, gas storage, sorption, separation, catalysis, non-linear optical properties, *etc* [1–6]. Intense research has resulted in a great number of coordination polymers with new structures and interesting luminescence and magnetic properties. In particular, lanthanide ions are interlinked by organic H<sub>3</sub>ssa and 2-stp and/or carboxylate groups to form many diversity structures [6–9]. Moreover, these ions with special luminescent resulting from 4f electrons, which illustrate the coordination polymers are intriguing and remarkable suitable for the development of technological devices [10]. Furthermore, Eu(III) compounds shows strong yellow luminescence in the visible region, making them to be good candidates for fluoroimmunoassays and structural and biological probes [11,12]. Due to the lanthanide ions have poor absorption coefficient, the emission intensities are very low. Using the suitable chromophore (antenna ligand or sensitizer), which is absorbed as a strongly in the UV-visible region and transfer its energy to the excited states Ln(III) ions. Therefore, the choice of appropriate ligand is very important [13,14].

Aromatic carboxylic ligands are good luminescent chromophores, resulting in the intense luminescence of lanthanide ions. An important example of a chromophoric carboxylate ligand is the 5-sulfosalicylic acid (H<sub>3</sub>ssa). The choice of 5-sulfosalicylic acid (H<sub>3</sub>ssa) can be attributed to the following properties: 5-sulfosalicylic acid (H<sub>3</sub>ssa) has interesting biological activities, such as anti-microbial, anti-fungal and anti-inflammatory activities. The H<sub>3</sub>ssa ligand involves three functional groups, –SO<sub>3</sub>H, –COOH and –OH, which can be partly or fully deprotonated in different forms, fabricating many novel architectures [15].

In this work, Eu<sup>III</sup> ion is aimed to gain luminescence properties by antenna effect. After ligand-mediated excitation of the complex, it shows the characteristic visible and/or near infrared (NIR) luminescence of the corresponding Ln<sup>III</sup> ion. This is attributed to efficient energy transfer from the ligands to the central Ln<sup>III</sup> ion (antenna effect). In this context, we report the synthesis, structural characterization and photophysical properties of novel Lanthanide-salicylate organometallic complex, {[Eu(ssa)(H<sub>2</sub>O)<sub>2</sub>].H<sub>2</sub>O}<sub>n</sub> (**1**). The crystal structure of complex **1** were previously reported at room temperature [16,17]. Herein, the crystal structure of complex **1** have been re-determined by the single crystal X-ray diffraction analysis and by FT-IR, UV-

Vis and photoluminescence properties this complex will report in this work .

## 2. Material and method

### 2.1. Materials and measurements

All chemical reagents and solvents were purchased from Aldrich and used without further purification. Elemental (C, H, N) analyses were carried out by standard methods with a LECO, CHNS-932 analyzer. FT-IR spectra were measured with a Perkin-Elmer Spectrum 65 instrument in the range of 4000 – 600 cm<sup>-1</sup>. Powder X-ray measurements were performed using Cu K $\alpha$  radiation ( $\lambda_{K\alpha}$  = 1.5418 Å) on a Bruker-AXS D8-Advance diffractometer equipped with a secondary monochromator. The data were collected in the range 5° < 2 $\theta$  < 60° in  $\theta$  –  $\theta$  mode with a step time of ns (5 s < n < 10 s) and step width of 0.02°. Solid state luminescence spectra in the visible and NIR region were measured at room temperature with an ANDOR SR500i-BL Photoluminescence Spectrometer, equipped with a triple grating and used an air-cooled CCD camera as detector. The measurements were done by using the excitation source (349 nm) of a Spectra-physics Nd:YLF laser with a 5 ns pulse width and 1.3 mJ of energy per pulse as the source.

### 2.2. Preparation of {[Eu(ssa<sup>-</sup>)(H<sub>2</sub>O)<sub>2</sub>](H<sub>2</sub>O)}<sub>n</sub>(**1**)

Complex **1** were prepared under hydrothermal condition. A mixture of Eu(NO<sub>3</sub>)<sub>3</sub>·(H<sub>2</sub>O)<sub>5</sub> (0.0428 g, 0.1 mmol) and H<sub>3</sub>ssa (0.0254 g, 0.1 mmol) in 20 ml of distilled water was sealed into a bomb equipped with a Teflon liner (45 ml) and then heated 140 °C for 5 days. The final pH values of these reactions media are close to 4.0. Crystals of complex **1** (yellow) was collected and washed distilled water. Anal. Calcd (%) for C<sub>7</sub>H<sub>9</sub>EuO<sub>9</sub>S: C, 19.96; H, 2.15; Found: C, 19.92; H, 2.14.

### 2.3. X-ray structure determination

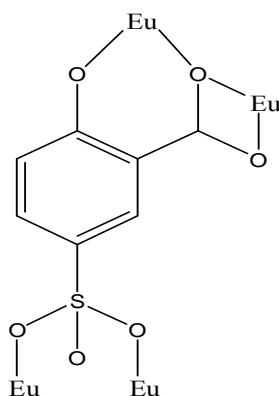
X-ray single crystal data for complex **1** were collected on a Xcalibur, Eos diffractometer using MoK $\alpha$  radiation at room temperature (293 K). The data were collected for Lorentz, polarization and absorption effects using the analytical numeric absorption correction technique [18]. Using OLEX2 [19], the structure was solved by direct methods using SHELXS [20] and refined by full-matrix least-squares based on |F<sub>obs</sub>|<sup>2</sup> using SHELXL [20]. The non-hydrogen atoms were refined anisotropically, while the hydrogen atoms, generated using idealized geometry, were made to “ride” on their parent atoms and used in the structure factor calculations. Detail of the supramolecular  $\pi$ -interactions were calculated PLATON 1.17 program [21].

### 3. Result and discussion

#### 3.1. Structural description of complex 1

The crystal data and structure refinement details for complex **1** are listed in Table 1. Selected bond lengths and angles are summarized in Table 2. Possible hydrogen bond geometry ( $\text{\AA}$ ,  $^\circ$ ) and distance between ring centroids [ $\text{\AA}$ ] are also given in Table 3. Representative structural diagrams of complex **1** are shown in Figure 1, while packing diagrams are given in Figures 2, 3 and 4. The X-ray crystal structure analysis reveals that complex **1** consists of a Eu(III) ion, one  $\text{ssa}^-$  ligand, two coordinated water molecules and one free water molecule in the asymmetric unit. The eight-coordinated environment around each Eu(III) centre is completed by six oxygen atoms from four  $\text{ssa}^-$  ligands and two oxygen atoms from a coordinated water molecule with Eu–O bond lengths in the range of 2.257(10) – 2.657(10)  $\text{\AA}$  (Figure 1). All bond distances and angles are comparable to similar structures [16,17,22].

The  $\text{ssa}^-$  ligand serves as  $\mu_4$ -bridge linking four Eu(III) ions by a bidentate bridging mode of the sulfonate, a bidentate chelating-bridging mode of the carboxylate, and a monodentate mode of hydroxyl group (see Scheme 1). Eu(III) ions are bridged by oxygen atoms of carboxyl and sulfonate groups, forming 1D chains. The neighbouring Eu...Eu distance is 4.654  $\text{\AA}$ . The adjacent phenyl rings of  $\text{ssa}^-$  ligands are parallel with the vertical distance of 3.71  $\text{\AA}$ , which shows there exist strong  $\pi$ - $\pi$  interactions in the 1D chains (Figure 2). Moreover, the 1D chains are linked together with ligand  $\text{ssa}^-$  to form a 2D wave-like layers (Figure 3). All these wave-like layers are further interlinked via strong hydrogen bonding interaction resulting in 3D architecture (Figure 4).



Scheme 1. The coordination mode of  $\text{ssa}^-$  ligand of complex **1**.

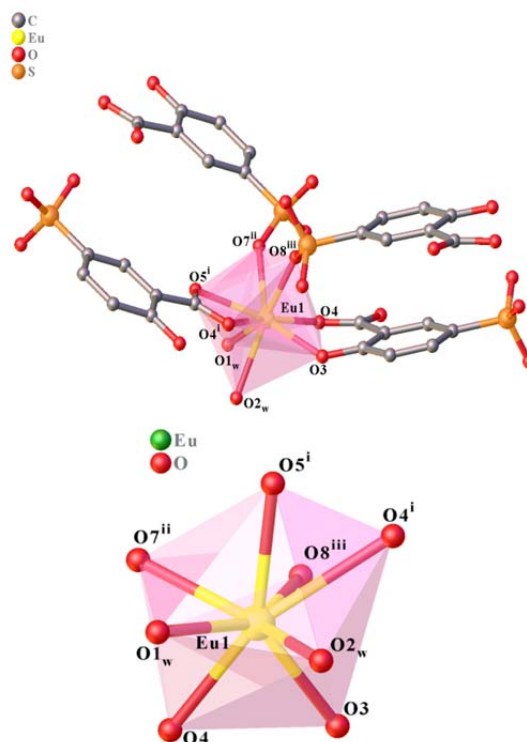


Figure 1. (a) Perspective view of the coordination environment of Eu(III) ions of complex **1**. Lattice water molecule and hydrogen atoms are omitted for clarity. (b) the trigonal dodecahedron coordination polyhedron of Eu(III) atom in complex **1**.

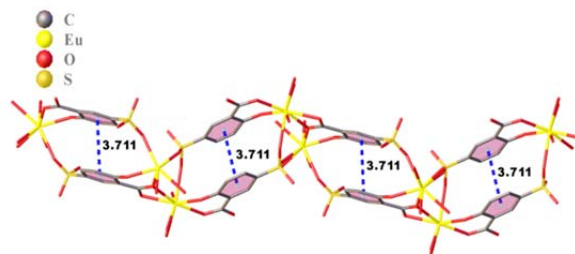


Figure 2. A view of  $\pi$ - $\pi$  interaction in the 1D infinite chains.

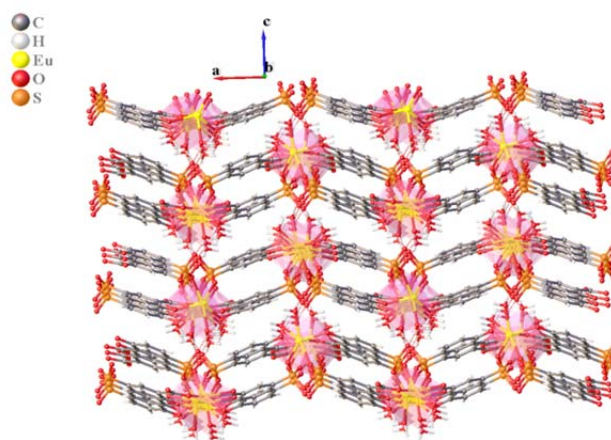


Figure 3. The 2D wave-like layers in the  $ab$ -plane.

Table 1. Crystallographic data and refinement parameters of compound **1**.

Eu(1)	
Chemical Formula	C <sub>7</sub> H <sub>9</sub> EuO <sub>9</sub> S
M <sub>w</sub>	421.16
Crystal System	Orthorhombic
Space Group	Pbcn
a/Å	15.8832(13)
b/Å	9.1963(6)
c/Å	14.9987(11)
α, β, γ	90, 90, 90
V/Å <sup>3</sup>	2190.8(3)
T/K	293(2)
Z	8
ρ <sub>calc</sub> /g.cm <sup>-3</sup>	2.554
μ/mm <sup>-1</sup>	5.956
Reflections Collected	5878
Independent Reflections	2232 (R <sub>int</sub> =0.068)
R <sub>1</sub> [I>2σ(I)]	0.0709
wR <sub>2</sub> (all data)	0.1980
S	1.015

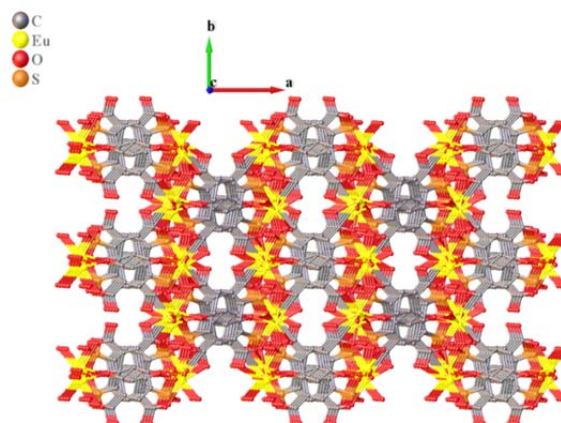


Figure 4. A view of 3D framework stabilized by hydrogen bonding interactions between different layers in *ac* plane (Hydrogen atoms have been omitted for clarity).

Table 2. Selected bond length (Å), and bond angles (°).

Eu1 – O1	2.397(9)	Eu1 – O4 <sup>i</sup>	2.657(10)
Eu1 – O2	2.388(10)	Eu1 – O5 <sup>i</sup>	2.426(10)
Eu1 – O3	2.257(10)	Eu1 – O7 <sup>ii</sup>	2.400(9)
Eu1 – O4	2.373(10)	Eu1 – O8 <sup>iii</sup>	2.377(10)
O1 – Eu1 – O2	69.6(3)	O3 – Eu1 – O4 <sup>i</sup>	87.6(3)
O1 – Eu1 – O3	125.3(4)	O3 – Eu1 – O5 <sup>i</sup>	138.3(4)
O1 – Eu1 – O4	76.2(3)	O3 – Eu1 – O7 <sup>ii</sup>	131.9(3)
O1 – Eu1 – O4 <sup>i</sup>	120.5(3)	O3 – Eu1 – O8 <sup>iii</sup>	77.4(4)
O1 – Eu1 – O5 <sup>i</sup>	84.0(4)	O4 – Eu1 – O4 <sup>i</sup>	159.55(17)
O1 – Eu1 – O7 <sup>ii</sup>	76.4(3)	O4 – Eu1 – O5 <sup>i</sup>	149.0(4)
O1 – Eu1 – O8 <sup>iii</sup>	151.7(3)	O4 – Eu1 – O7 <sup>ii</sup>	73.3(3)
O2 – Eu1 – O3	79.8(4)	O4 – Eu1 – O8 <sup>iii</sup>	98.5(3)
O2 – Eu1 – O4	109.0(4)	O4 <sup>i</sup> – Eu1 – O5 <sup>i</sup>	50.7(3)
O2 – Eu1 – O4 <sup>i</sup>	70.3(3)	O4 <sup>i</sup> – Eu1 – O7 <sup>ii</sup>	120.0(3)
O2 – Eu1 – O5 <sup>i</sup>	85.4(4)	O4 <sup>i</sup> – Eu1 – O8 <sup>iii</sup>	72.3(3)
O2 – Eu1 – O7 <sup>ii</sup>	143.9(4)	O5 <sup>i</sup> – Eu1 – O7 <sup>ii</sup>	79.1(4)
O2 – Eu1 – O8 <sup>iii</sup>	136.6(3)	O5 <sup>i</sup> – Eu1 – O8 <sup>iii</sup>	87.7(4)
O3 – Eu1 – O4	72.4(3)	O7 <sup>ii</sup> – Eu1 – O8 <sup>iii</sup>	75.4(3)

Symmetry transformations used to generate equivalent atoms for complex 1: (i) 1/2-x, 1/2+y, +z, (ii) -1/2+x, 1/2-y, 1-z, (iii) 1-x, 1-y, 1-z.

Finally, before proceeding to the photoluminescence characterization, we note that PXRD patterns for bulk microcrystalline sample of **1** were consistent with the exclusive presence of the phase identified in the single crystal experiment (Figure 5).

Table 3. Hydrogen bond geometry (Å, °) and distance between ring centroids (Å) for complex **1**.

D–H...A*	D–H	H...A	D...A	D–H...A	Symmetry Code
O1–H1A...O9	0.90	2.37	2.93(2)	120	
O1–H1B...O6	0.90	2.28	3.130(13)	156	1-x,y,3/2-z
O2–H2A...O6	0.89	2.05	2.913(15)	165	1-x,y,3/2-z
O2–H2B...O9	0.89	2.46	2.97(2)	118	
O9–H9A...O2	0.85	2.52	2.97(2)	115	
O9–H9B...O1	0.85	2.38	2.93(2)	123	
<b>Cg(I)...Cg(J)</b>					
Cg(1)...Cg(1)			4.167(7)		1-x,y,3/2-z
Cg(1)...Cg(1)			3.711(7)		1-x,1-y,1-z

D: Donor, A: Acceptor, Cg(I): Plane number I (=ring number in () above), Cg(1): C2-C3-C4- C5-C6-C7.

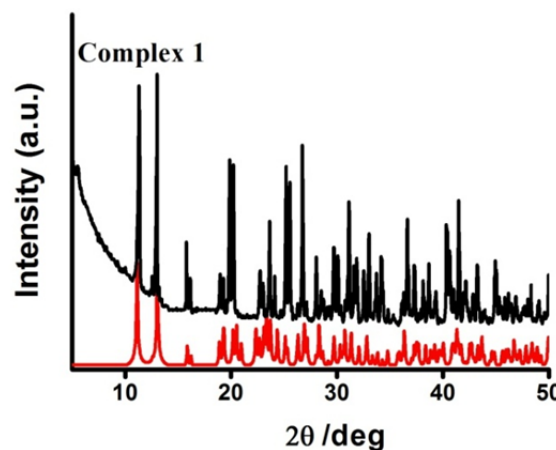


Figure 5. X-ray powder diffraction pattern of complex **1** (Red-simulated CIF, Black-experimental)

### 3.2. FT-IR spectra and UV-Vis analysis

The FT-IR and solid state UV-Vis spectra of complex **1** were shown in comparison with that of its free ligand H<sub>3</sub>ssa (Figures 6 and 7). As seen from Figure 6, the broad peak at 3337 cm<sup>-1</sup> indicates the presence of  $\nu(\text{O-H})$  vibration associated with coordinated water molecules [23]. The carboxyl band of the free ligand H<sub>3</sub>ssa is found at 1657 cm<sup>-1</sup>. The absence of this band indicates that H<sub>3</sub>ssa are completely deprotonated in the form of ssa<sup>-</sup> anions upon reaction with the metal ions [10]. For complex **1**, two medium bands are at 1606 and 1379 cm<sup>-1</sup> corresponding to the stretching vibrations of asymmetric and symmetric carboxyl groups. The difference of asymmetric and symmetric vibration bands is about 227 cm<sup>-1</sup>, which indicating carboxylic groups are coordinated to Eu<sup>III</sup> in chelating mode [23]. The two peaks at 1468 and 1434 cm<sup>-1</sup> were assigned to the skeletal vibration of benzene ring bonds. The band at 1330 and 1255 cm<sup>-1</sup> resulted from asymmetric stretching vibration, whereas the bands at 1139 and 1043 cm<sup>-1</sup> were caused by the symmetric stretching vibrations of the sulfonate groups in ssa<sup>-</sup> [7]. The medium bands at 911 and 832 cm<sup>-1</sup> are ascribed to O-Eu-O stretching and the strong band at 670 cm<sup>-1</sup> and the weak band at 809 cm<sup>-1</sup> are characteristic of Eu-O stretching vibration [24]. To obtained results are in agreement with data from the single crystal X-ray analysis.

Figure 7 show the UV-vis spectra of the free ligand (H<sub>3</sub>ssa) and its Eu(III) complex were recorded in solid state. In the absorption spectra of free ligand, there are two set of broad bands. The high energy band maxima at 306 nm (32680 cm<sup>-1</sup>) and the other absorption band maxima at 399 nm (25062 cm<sup>-1</sup>) for (H<sub>3</sub>ssa) ligand. The absorption spectra of complex **1** indicated different absorption pattern as compared to its free ligand H<sub>3</sub>ssa. The dominated absorption band observed with the maxima at 426 nm (23474 cm<sup>-1</sup>) in the spectrum of complex **1**, could be assigned to  $\pi-\pi^*$  or  $n-\pi^*$  transition of the ligand H<sub>3</sub>ssa. The shifting of absorption band in the spectra of the complex **1** signify the H<sub>3</sub>ssa ligand coordination with lanthanide ion [25–27].

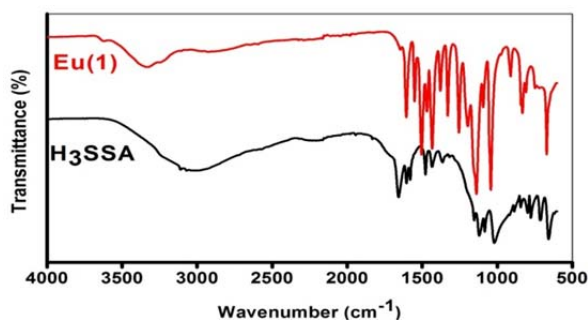


Figure 6. IR spectra of the free ligand H<sub>3</sub>ssa and its complex.

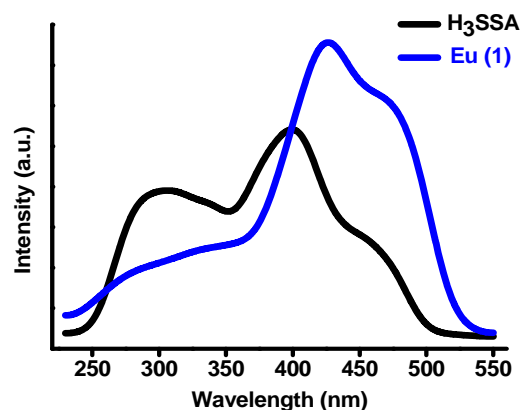


Figure 7. Solid state UV-Vis spectra of the free ligand H<sub>3</sub>ssa and its complex.

### 3.5. Photoluminescence Properties

The solid state luminescent properties of the free ligand H<sub>3</sub>ssa and complex **1** were investigated at room temperature in the visible region upon excitation at  $\lambda_{\text{exc.}} = 349$  nm (Figure 8). The free ligand H<sub>3</sub>ssa shows broad emission band at  $\lambda_{\text{max}} = 465$  nm in the blue-region, which may be assigned to the  $n \rightarrow \pi^*$  or  $\pi \rightarrow \pi^*$  electronic transition intraligand charge transfer (ILCT) [28].

Complex **1** exhibit six characteristic peaks: 523 nm correspond to  ${}^5\text{D}_1 \rightarrow {}^7\text{F}_0$ , 577, 590, 618, 652 and 700 nm originating from the transition of  ${}^5\text{D}_0 \rightarrow {}^7\text{F}_J$  ( $J = 0-4$ ) of the Eu<sup>III</sup> ion, respectively [29–32]. The appearance of the symmetry-forbidden emission band at 577 nm ( ${}^5\text{D}_0 \rightarrow {}^7\text{F}_0$ ) indicates that the Eu<sup>III</sup> ions in complex **1** possess the noncentro-symmetric coordination environment [33]. It is well known that the electric dipole  ${}^5\text{D}_0 \rightarrow {}^7\text{F}_2$  transition is the most intense emission which leads to yellow luminescence, and hypersensitive transition, whose intensity is very sensitive to site symmetry, while the magnetic dipole transition  ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$  is insensitive and hardly varies with the coordination environment [34,35]. The intensity ratio  $I({}^5\text{D}_0 \rightarrow {}^7\text{F}_2) / I({}^5\text{D}_0 \rightarrow {}^7\text{F}_1)$  is equal to ca. 12, indicating that the Eu(III) ions occupy non-inversion symmetry sites, which is in accordance with the results of its single crystal structural [36,37]. On the other hand, the emission spectrum of complex **1** indicates also a broad emission band around 355 to 495 nm with the maximum wavelength of 410 nm. In other words, complex **1** show a broad band, the inefficient quenching of the ligand-centred emission can be further seen from broad bands with low intensity. The spectrum indicates that the energy transfer from the ligand to Eu<sup>III</sup> ion is moderately efficient [9].

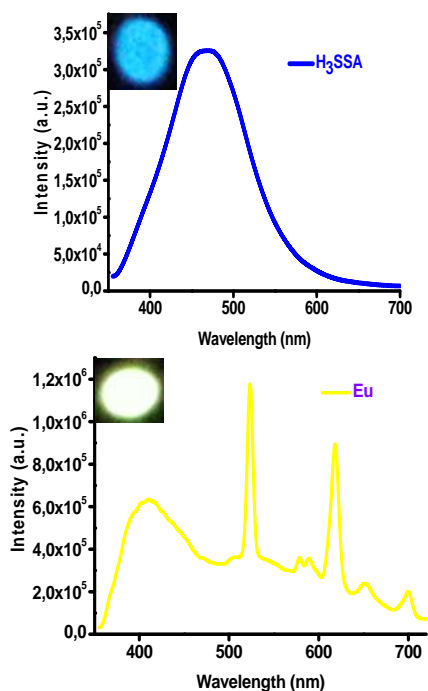


Figure 8. (a) The solid state emission spectrum of the free ligand H<sub>3</sub>ssa and (b) complex **1** in the UV-Visible region at room temperature ( $\lambda_{exc.} = 349$  nm).

Generally, the lanthanide-based organic complexes, due to the intraconfigurational  $f \leftarrow f$  transition are parity-forbidden and also Ln ions have low absorption coefficient. Therefore, researchers have proposed the use of organic chromophores in the lanthanide compounds which absorb light in the UV region and transfer the excitation energy from the chromophore to central lanthanide ions via an antenna effect (Figure 9), which can be seen in spectrum as the overlaps between the excitation spectrum of the complex and the absorption spectra of its free ligands (Figure 10) [38–40]. In briefly, Lanthanide-sensitization processing consists of three steps (Figure 9), (i) the ligands absorb the energy and are excited to the singlet (S) excited state, (ii) the energy of the (S) excited state is transferred to its triplet (T) level via intersystem crossing (ISC) and (iii) the energy is transferred to the 4f levels of the Ln(III) ions, resulting in the characteristic emissions of the sensitized Ln(III) ions [40].

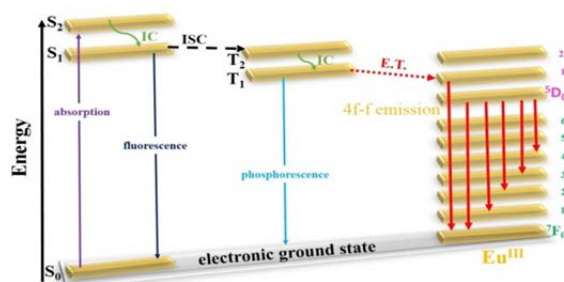


Figure 9. Jablonski diagram showing basic photophysical process (left) and energy level diagrams of Eu<sup>III</sup> ion (right).

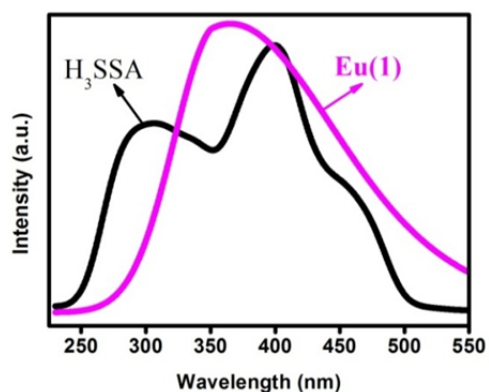


Figure 10. The excitation spectrum for complex **1** and absorption spectra of the free H<sub>3</sub>ssa ligand as solid state.

#### 4. Conclusion

The new Eu(III) based coordination polymer has been synthesized successfully under hydrothermal condition with using 5-sulfosalicylic acid (H<sub>3</sub>ssa) ligand. The crystal structure of complex **1** has been re-determined by single x-ray diffraction analysis. And also complex characterized by FTIR, UV-Vis spectra, elemental analysis and also as a complementary characterization PXRD. Structural analysis of complex **1** exhibits that each Eu(III) ions is eight coordinated by oxygen atoms of ssa ligands and water molecules. The solid state photoluminescence properties have been investigated at room temperature in UV-Visible region for free H<sub>3</sub>ssa ligand and its complex **1**. The photoluminescence measurement shows yellow emission for complex **1**, which is attributed to the f-f electronic transitions. Moreover, complex **1** shows a broad band, the inefficient quenching of the ligand-centred emission can be further seen from the broad bands with low intensity around 355-495 nm. The emission spectrum shows that the energy transfer from ligands to complex **1** is moderately efficient. In generally, the coordinated water molecules can quench the emission of Ln(III) ion, but complex **1** indicates intense emission, which also may be accounted for by the sensitization of the ligands and the extensive  $\pi$  system of the polymer structures.

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### Supplementary Material

Crystallographic data for the structural analysis have been deposited with the Cambridge Crystallographic Data Centre, CCDC No. 1565766 for complex **1**. Copies of the data can be obtained free of charge from the Director, CCDC, 12 Union Road, Cambridge CB2 1EZ, UK (fax: +44-1223-336-033; e-mail: deposit@ccdc.cam.ac.uk or www: <http://www.ccdc.cam.ac.uk>).